

SINGLE REPLACEMENT, COMBUSTION, SYNTHESIS, AND DECOMPOSITION, REACTIONS

Write the balanced chemical equation for each of the following reactions. Be sure to include states of matter:

Single replacement practice – be sure to check your activity series!:

1. Magnesium is mixed with Silver Nitrate solution. Write the balanced chemical equation.
2. Copper is combined with Zinc acetate solution. Write the balanced chemical equation.
(Copper forms a Copper (II) ion when reacted)
3. Zinc is mixed with a solution of Copper (II) Chloride. Write the balanced chemical equation

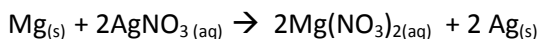
Mixed Review: Combustion, synthesis, decomposition, and single replacement. Write the balanced chemical equation for each.

4. A solution of Hydrogen peroxide (H_2O_2) decomposes into water and hydrogen gas.
5. Aluminum is placed in a solution of Iron (II) Nitrate.
6. Ethene (C_2H_6) is burned.
7. Ammonia (NH_3) is formed from a combination of its elements.
8. Chlorine gas is placed in a solution of sodium bromide.

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Single replacement practice – be sure to check your activity series!:

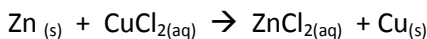
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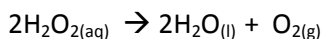
No Reaction – Zinc is higher than Copper on the activity series

3. Zinc is mixed with a solution of Copper (II) Chloride. Write the balanced chemical equation

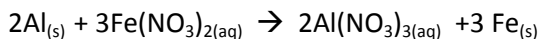


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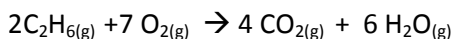
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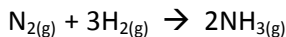
5. Aluminum is placed in a solution of Iron (II) Nitrate.



6. Ethene (C_2H_6) is burned.



7. Ammonia (NH_3) is formed from a combination of its elements.



8. Chlorine gas is placed in a solution of sodium bromide.

